pH Notes 11.16

Bio ING/HON

**Dissociation of Water**

**1. define:**

**2. illustrate:**

pH Scale: Label Acids, Neutral, and Bases (Alkaline)

🡨---------------------------------------X-----------------------------------🡪

pH Measures the Concentration of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

More \_\_\_\_\_ 🡪 More \_\_\_\_\_\_ 🡪 Lower pH

More \_\_\_\_\_ 🡪 More \_\_\_\_\_\_ 🡪 Higher pH

ACID: A substance which \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ when dissolved in water.

BASE: A substance which \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ when dissolved in water.

**Acids Basics ;)**

|  |  |
| --- | --- |
| **Acids** | **Bases** |
|  |  |
|  |  |
|  |  |
|  |  |

A change of 1 unit on the pH scale indicates a \_\_\_\_\_\_\_\_\_\_\_ change in ion concentration!

Three ways to say the same thing: what makes an acid? \_\_\_\_\_\_\_\_\_\_\_\_\_, aka \_\_\_\_\_\_\_\_\_\_\_\_\_, aka \_\_\_\_\_\_\_\_\_\_\_\_\_\_

* **\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_** = H3O+. Higher concentration 🡪 more acidic solution 🡪 lower pH
* **\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_** = -OH. Higher concentration 🡪 more basic solution 🡪 higher pH